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### **Advance Study Assignment Experiment 30**

Experiment 30 Advance Study Assignment: Determination of Iron by Reaction with Permanganate-A Redox Titration Section Name  
son in acidic solution 1. Write the balanced net ionic equation for the reaction between  $\text{MnO}_4^-$  ion and  $\text{Fe}^{2+}$   
 $\text{MnO}_4^- + 5 \text{Fe}^{2+} + 8 \text{H}^+ \rightarrow \text{Mn}^{2+} + 5 \text{Fe}^{3+} + 4 \text{H}_2\text{O}$

**Solved: Experiment 30 Advance Study Assignment:**

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## **Determinat ...**

Section Experiment 30 Advance Study Assignment:

Determination of Iron by Reaction with Permanganate Redox Titration

1. Write the balanced net ionic equation for the reaction between  $\text{MnO}_4^-$  ion and  $\text{Fe}^{2+}$  ion in acidic solution. 2. How many moles of  $\text{Fe}^{2+}$  ion can be oxidized by  $1.4 \times 10^2$  moles  $\text{MnO}_4^-$  ion in the reaction in Question 1? 3.

## **Solved: Section Experiment 30 Advance Study Assignment: De ...**

Experiment 30 Advance Study Assignment: Determination of Iron by Reaction with Permanganate-A Redox Titration 1. Write the balanced net ionic equation for the reaction between  $\text{MnO}_4^-$  ion and  $\text{Fe}^{2+}$  ion in acidic solution. 2. How many moles of  $\text{Fe}^{2+}$  ion can be oxidized by  $1.4 \times 10^2$  moles  $\text{MnO}_4^-$  ion in the reaction in Question 1? 3.

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### **Solved: Experiment 30 Advance Study Assignment: Determinat ...**

Question: Name Section Experiment 30 Advance Study Assignment: Determination Of Iron By Reaction With Permanganate-A Redox Titration 1. Write The Balanced Net Ionic Equation For The Reaction Between  $MnO_4^-$  Ion And  $Fe^{2+}$  Ion In Acidic Solution. 2. How Many Moles Of The Ion Can Be Oxidized By  $1.4 \times 10^{-2}$  Moles  $MnO_4^-$  In The Reactice In Question ?

### **Name Section Experiment 30 Advance Study Assignmen ...**

Name Section Experiment 30 Advance Study Assignment: Determination of Iron by Reaction with 1. Write the balanced net ionic equation for the reaction between  $MnO_4^-$  ion and  $Fe^{2+}$  ion in acidic solution. Permanganate-A Redox Titration 2. How many moles of  $Fe^{2+}$  can be oxidized by  $1.4 \times 10^{-2}$  moles  $MnO_4^-$  in the reaction in Question 1? moles 3.

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in Question 1? moles. 3.

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