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KCl solution to
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to g KCl (2.95 g

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... $M_1V_1 = M_2V_2$.

In this problem,
the initial

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molarity is 3.00

M, the initial
volume is 2.50

mL or 2.50×10

-3 L and the

final volume is

0.175 L. Use

these known

values to

calculate the

final molarity,

M_2 : So, the

final

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16. For a solution to form, one substance must dissolve in another. For this to happen, the solute and solvent particles must ATTRACT ONE ANOTHER. 17.

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solution, energy
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$$\begin{aligned} 0.5 \text{ } ^\circ\text{C} &= ((4.68 \\ &^\circ\text{C kg mol}^{-1}) \times \\ &(0.60 \text{ g / MM})) / \\ 0.1 \text{ kg. MM} &= \\ (4.68 \times 0.60) / \\ (0.5 \times 0.1) &= 56 \\ \text{or } 6 \times 101. &\text{ an} \end{aligned}$$

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alternate solution for (c)
molality = 0.5

$$^{\circ}\text{C} / (4.68 \text{ } 0.5$$

$$^{\circ}\text{C/m}) = 0.107$$

m. mol solute = (
0.107 mol / kg
solvent) x 0.100
kg solvent =

$$0.0107 \text{ mol. MM} =$$

$$0.60 \text{ g} / 0.0107$$

mol = 56 or 6 x
101. d) one

point.

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High. (C1) (V1) =

(C2) (V2)

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2 (V2 (concentrati
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The amount of a
substance
dissolved in a

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given amount of solvent is the concentration of the solute, which can be expressed in terms of molarity or molality. If you know the molarity of a solution, you can determine the exact volume

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of the solution
that contains a
desired amount
of the solute.

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a large solute
to solvent
ratio. F. What

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add more solvent
to the solution,
changing only
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Answer to: A
35.0 mL sample
of an HCl

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solution is placed in a flask with a few drops of phenolphthalein indicator. If it takes 38.2 mL of a 0.480 M...

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