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CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the

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percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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Answer Key CHAPTER 9 REVIEW
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ANSWER Answer the following questions
in the space provided. 1. Given the
following equation: $C_3H_4(g) + xO_2(g)$
 $? 3CO_2(g) + 2H_2O(g)$ 4 a. What is the
value of the coefficient x in this
equation? 40.07 g/mol b.

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Describe a chemical reaction. Define reactant. Define product. Identify the products and reactants in a reaction. Identify a chemical change. Relate the symbols in a chemical equation to the words in a word equation. Write the word equation from a sentence.

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mass of CuO to 98.4 grams

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chemistry test / exam. Topics include the following: How To Convert Between Grams to ...

Stoichiometry Tutorial Review - Limiting Reagent - Chemistry Practice Problems

60.2 9 42.1 1 a. \ tt mash 01 ox aen Cas
i pridui.ed it 100. of lithium c a C ti. I o c.

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... The oxygen gas produced in part a has density of 1.43 g/L calculate the volume of this gas. 76
STOICHIOMETRY MODERN CHEMISTRY a.
... 81 g. A car air bag requires 70. L of nitrogen gas to inflate properly. The following ...

Date. FCHAPJ REV[EW.

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Stoichiometry Chapter 9 - Mrs. Baker's Chemistry Website

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class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.

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Chemistry Stoichiometry (Chapter 8)

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Chemistry 101: Chapter 2/3: Atoms,
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Chemistry Exam 1 Review;
Stoichiometry and Chemistry; MCAT
Chemistry Ch. 4 Compounds &
Stoichiometry; Macroeconomics
Exam/Midterm 2 KT; science 1150; Exam
1, RPTS 202; Chapter One and Chapter

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Six.

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The reaction stoichiometry problems in this chapter can be classified according to the information given in the problem and the information you are expected to find, the unknown. ... in chemical

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technology, chemistry, or other sciences. 290 Chapter 9

**CorrectionKey=NL-A DO NOT
EDIT--Changes must be made ...**

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CHAPTER 9 REVIEW Stoichiometry

SECTION 2 PROBLEMS Write the answer
on the line to the left Show all your work

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in the space provided 1 ____ The following equation represents a laboratory preparation for oxygen gas:
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles ...

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